PARTICLE PHYSICS AND COSMOLOGY

in the 20th century

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CERN

DISCLAIMER

This is a lecture about particle physics and cosmology.
It covers about 100 years of ideas, theories and experiments.

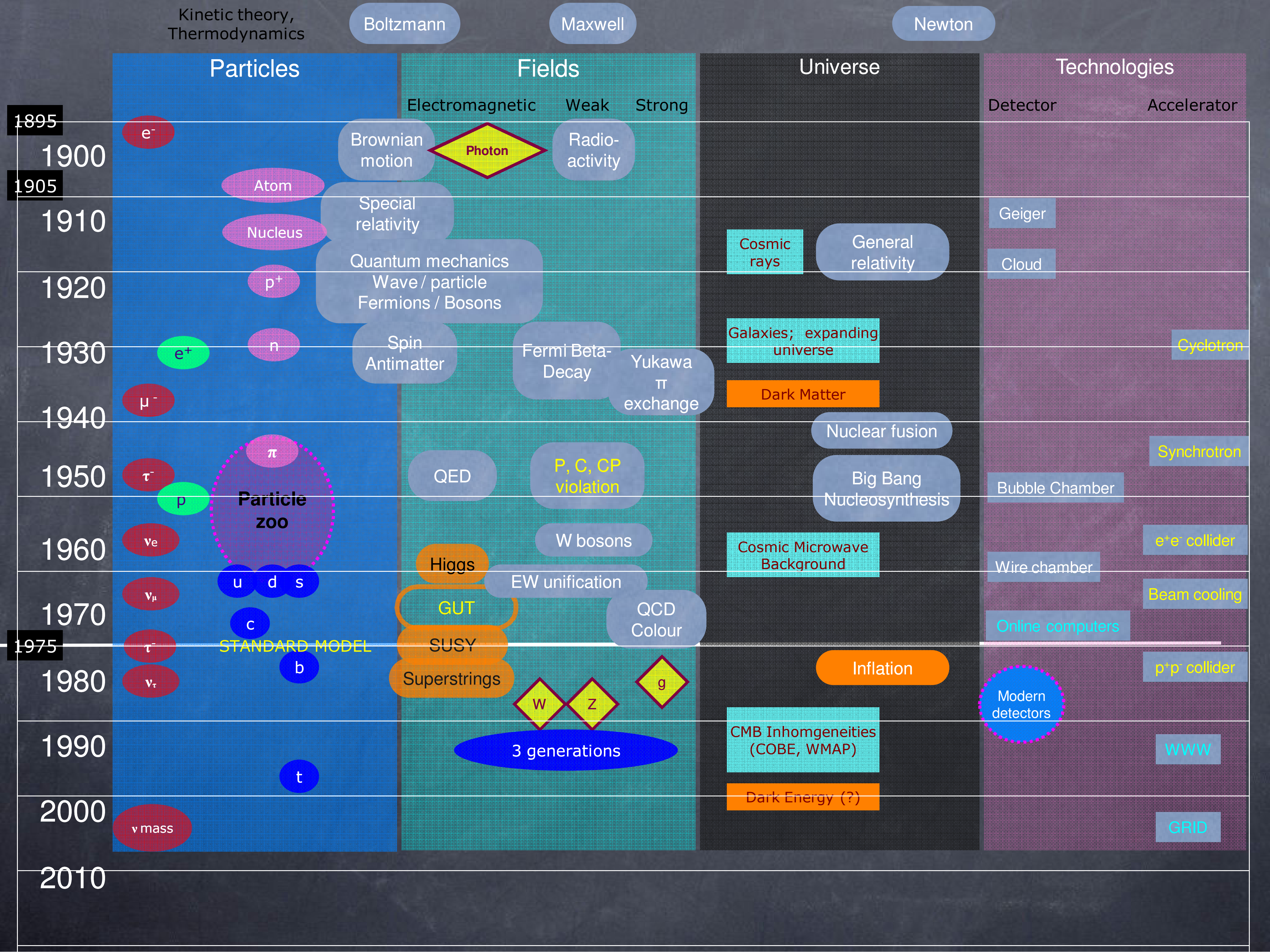
More than 50 Nobel prize winners on particle physics

Very difficult to be comprehensive, exact or in-depth

This is a broad overview about the main discoveries.

All that remains to do in physics is to fill in the sixth decimal place

(Albert Michelson, 1894)


Kinetic theory, Thermodynamics
Boltzmann
Maxwell
Newton
Particles
Fields
Electromagnetic
Weak
Strong
Universe
Technologies
Detector
Accelerator
1895
1900
1905
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1970
1975
1980
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e^-
Atom
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STANDARD MODEL
b
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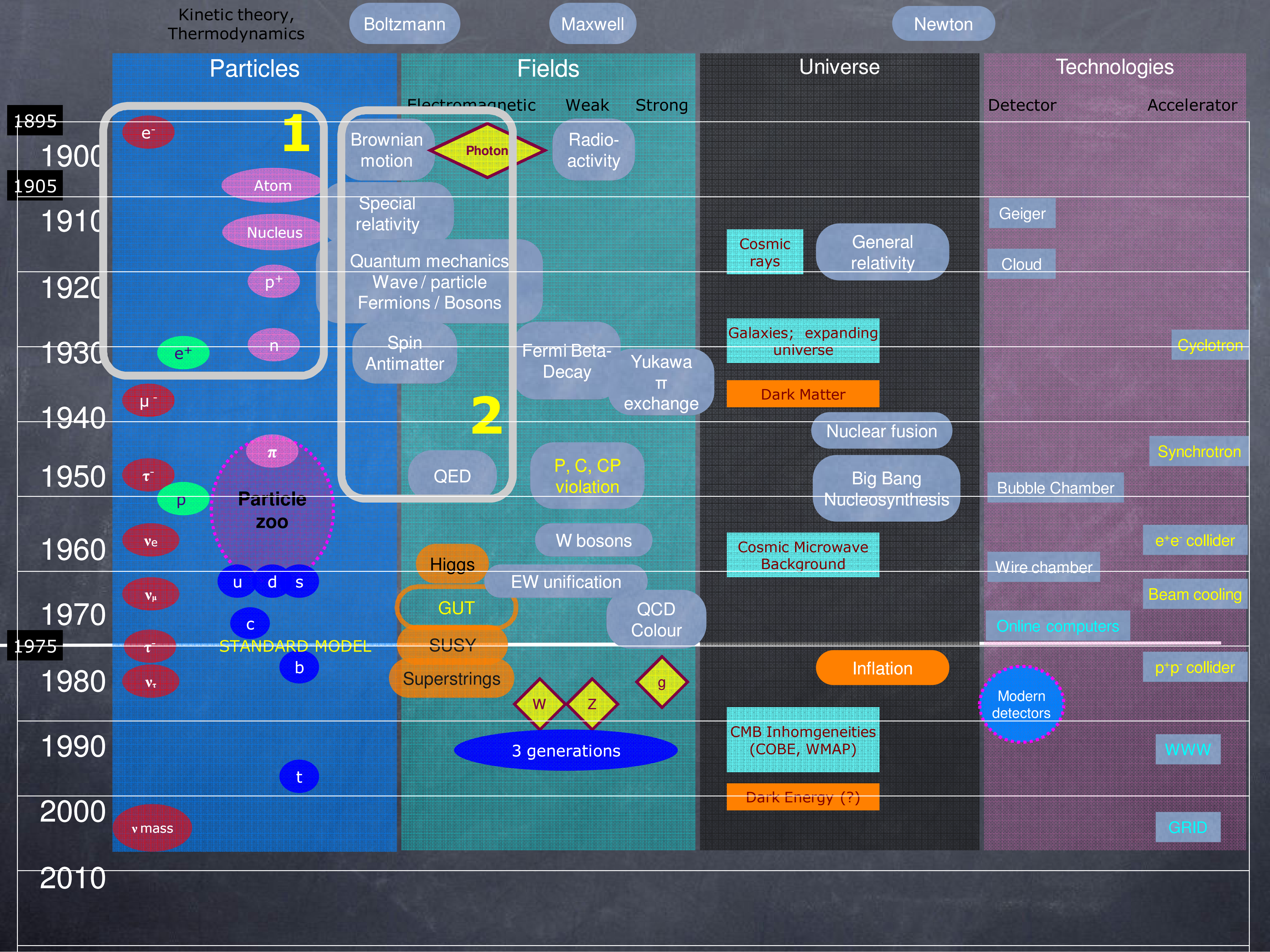
ν mass
Brownian motion
Photon
Radio-activity
Special relativity
Quantum mechanics
Wave / particle
Fermions / Bosons
Spin
Antimatter
Fermi Beta-Decay
Yukawa π exchange
QED
P, C, CP violation
W bosons
Higgs
EW unification
GUT
QCD Colour
SUSY
Superstrings
W
Z
g
3 generations
Cosmic rays
General relativity
Galaxies; expanding universe
Dark Matter
Nuclear fusion
Big Bang Nucleosynthesis
Cosmic Microwave Background
Inflation
CMB Inhomgeneities (COBE, WMAP)
Dark Energy (?)
Geiger
Cloud
Cyclotron
Synchrotron
Bubble Chamber
e^+e^- collider
Wire chamber
Beam cooling
Online computers
p^+p^- collider
Modern detectors
WWW
GRID



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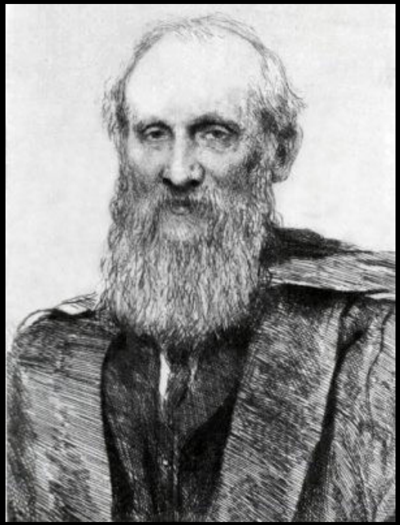
In the early 1900s, most physicists believed that physics was complete, described by classical mechanics, thermodynamics, and the Maxwell theory.



William Thomson
(Lord Kelvin)

Address to the British Association for the Advancement of Science, 1900

There is nothing new to be discovered in physics now, All that remains is more and more precise measurement.
(Lord Kelvin, 1900)

But Lord Kelvin also mentioned two ‘clouds’ on the horizon of physics:

1) Blackbody radiation

2) Michelson-Morley experiment

1900

Universe = solar system and the stars of our galaxy

Nobody knew how the sun produced its energy

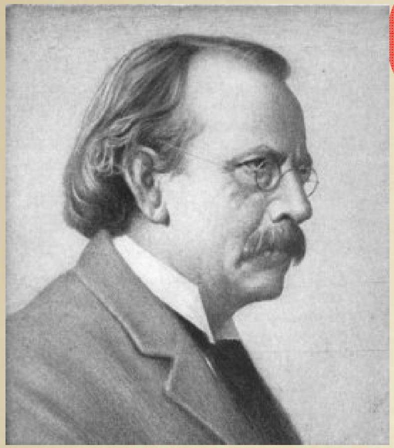
Nothing was known the structure of atoms and nuclei

Only two known fields: gravitation, electromagnetism

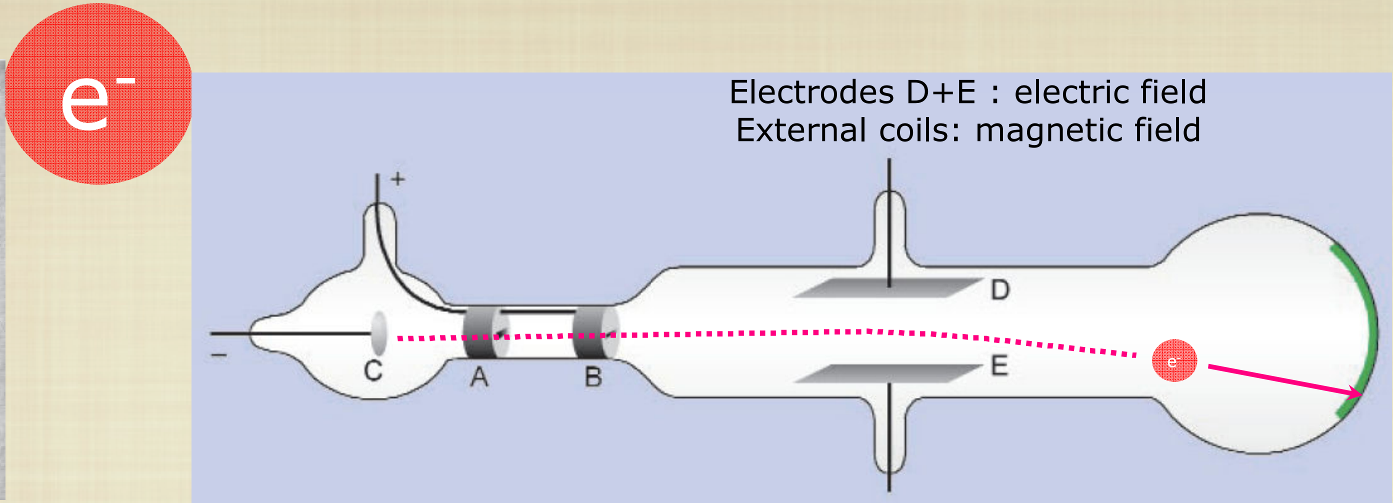
Nobody anticipated the incredible journey of physics in the next 100 years

PARTICLE SPECTRUM

1897



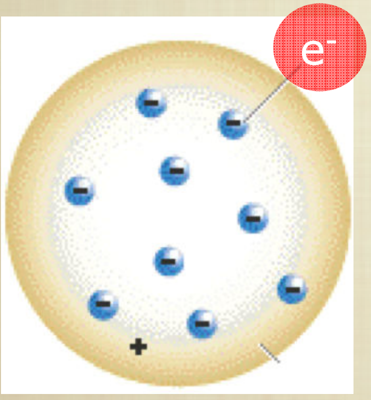
J.J. Thomson



Cathode ray experiments (~ TV)

'Rays' are charged corpuscles*
with unique charge/mass ratio

***later called 'electrons'**

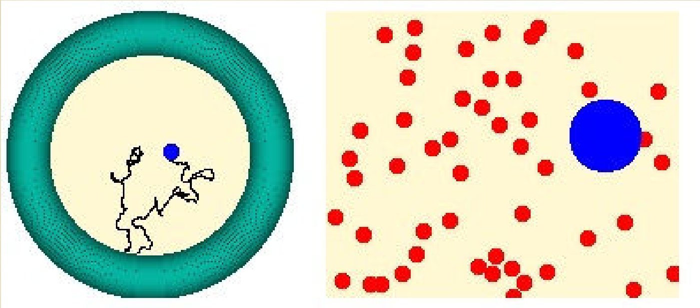


His 'plumpudding' model
of the atom (1904)

Electrons are sub-atomic particles!

PARTICLE SPECTRUM

Atom



Robert Brown (1827) observes random walk of small particles suspended in a fluid

Albert Einstein (1905) explains by kinetic theory that the motion is due to the bombardment by molecules

Francois Perrin (1907) uses Einstein's formula to confirm the theory and measure Avogadro's number

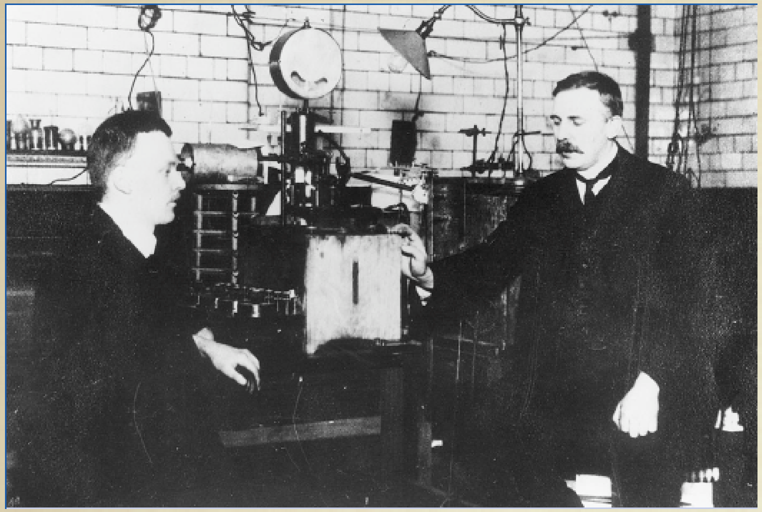
$$\langle x^2 \rangle = \frac{2kTt}{\alpha} = \frac{kTt}{3\pi\eta a}$$

The existence of atoms was proven

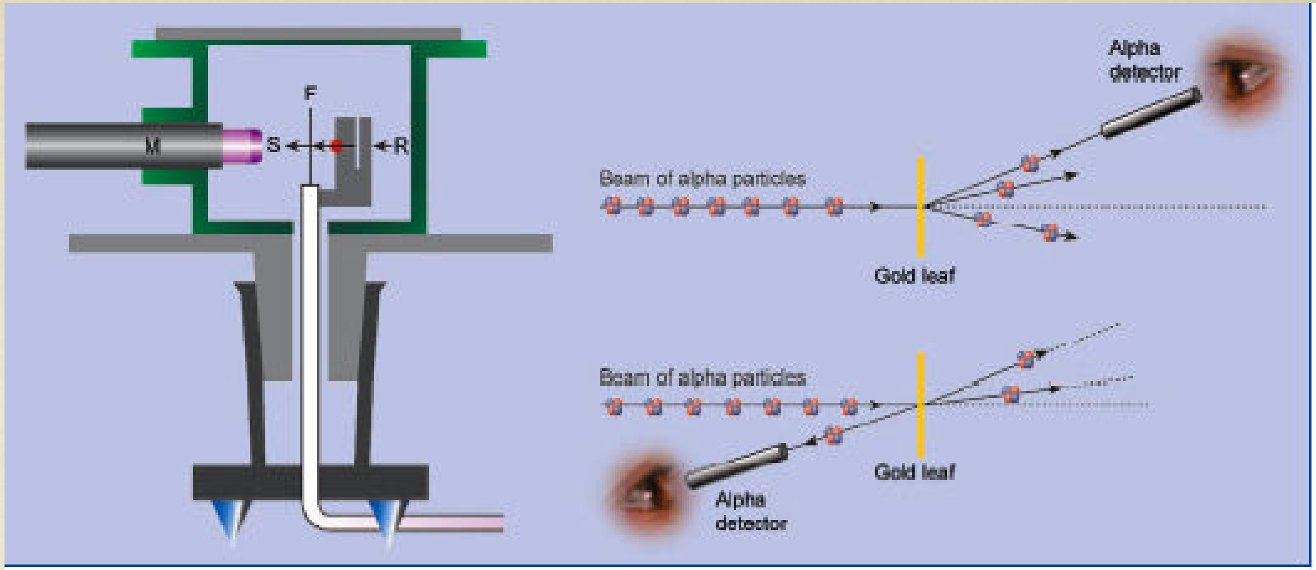
PARTICLE SPECTRUM

Nucleus

1911



Ernest Rutherford (r) and Hans Geiger (l) in Manchester



Geiger and Marsden fired alpha particles (He nuclei) on gold foils

1 in 8000 alpha particles were backscattered (> 90 deg)

This could not be explained by the 'plumpudding model'

Rutherford's explanation: all the mass of the atom is concentrated in the nucleus

Size: At minimum distance, Coulomb repulsion = kinetic energy: ~ 27×10^{-15} m (true value: 7.3)

Discovery of the nucleus

PARTICLE SPECTRUM

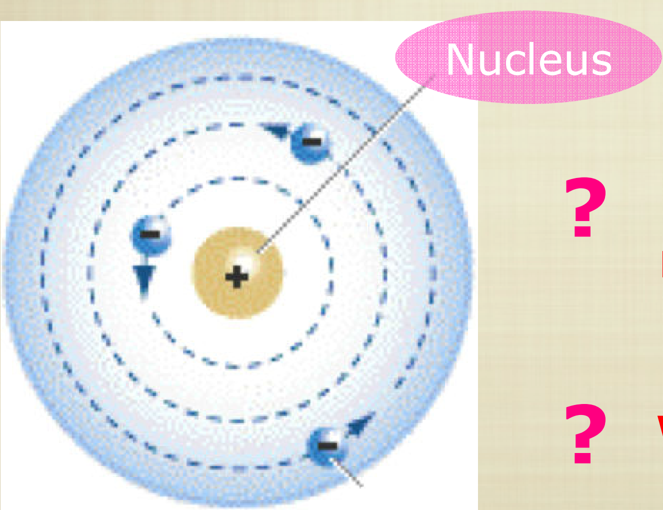
Nucleus

1911

Analogy with solar system:

If the nucleus had the size of the Sun

the electrons would orbit in 1000 x the distance of Sun-Earth



Rutherford's model
of the “empty” atom

? How can electrons orbit a nucleus without radiating their energy?

? What is the nucleus made of ?

PARTICLE SPECTRUM

1913

J. J. Balmer (1885) observed the emission spectrum of hydrogen

His empirical formula: $$\lambda = \frac{hm^2}{(m^2 - n^2)}$$

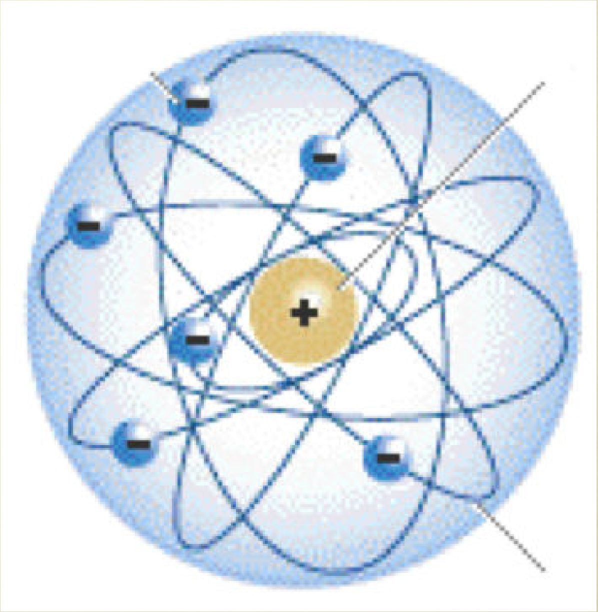
Niels Bohr visited Rutherford in 1913

he was the first to apply quantum ideas to atoms

- Quantization of angular momentum -> energy levels

$$\mathbf{L} = n \cdot \hbar = n \cdot \frac{h}{2\pi} \qquad E_n = \frac{-13.6\ \text{eV}}{n^2}$$

- Emission of radiation only during transitions
- Energy of photons = difference of energy levels



PARTICLE SPECTRUM

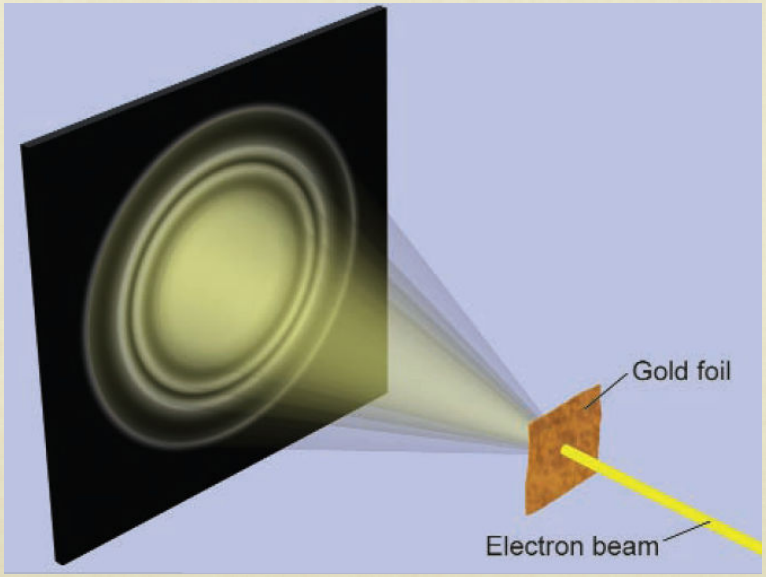
1923-1927

It took 10 more years to understand the mysterious rules governing the atomic world: quantum mechanics.

Louis de Broglie (1924)

Particles behave like waves

$$\lambda = \frac{h}{p}$$



*this hypothesis was confirmed in 1927
by electron diffraction (Davisson/Germer)

PARTICLE SPECTRUM

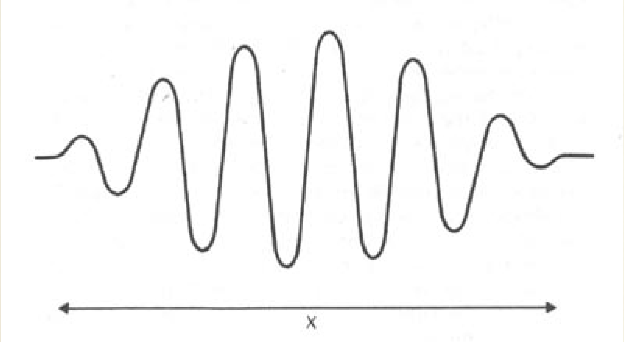
1923-1927

Heisenberg (1925)

Uncertainty relation

If particles are waves (of finite size), then there must be a limit to the precision of measurement between:

Position and momentum $\Delta x \Delta p \geq \frac{\hbar}{2}$



a wave packet corresponding to a particle located somewhere in the region X

Analogy:

Measurement time Δt of a signal leads to uncertainty of frequency (Fourier transform):

$\Delta f \, \Delta t \sim 1$

Energy and time $\Delta E \Delta t \geq \hbar$

PARTICLE SPECTRUM

1923-1927

Schrödinger
1926

Probability wave function

Excellent description
if v << c

If particles are waves -> describe by a wave equation

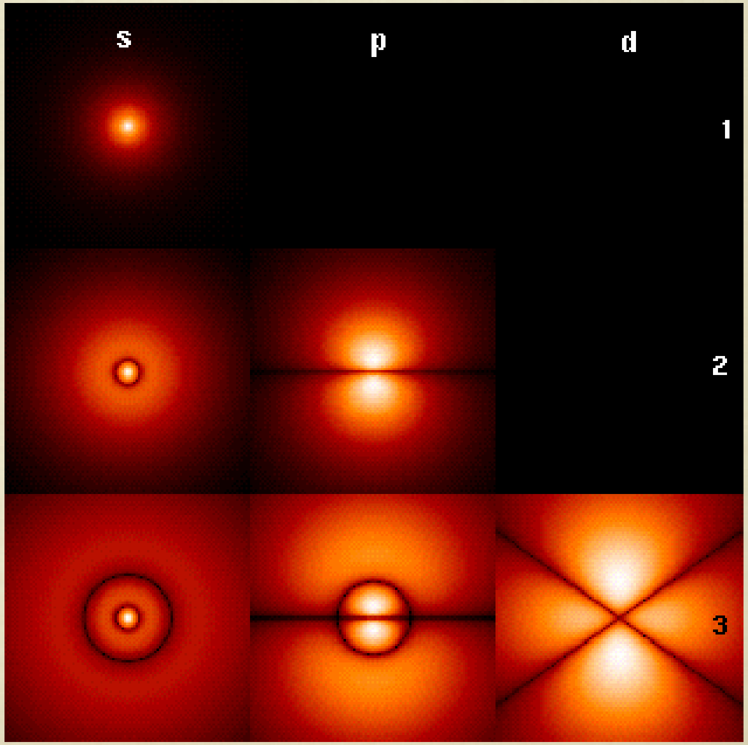
$$H\psi\left(\mathbf{r},t\right) = \left(T+V\right)\psi\left(\mathbf{r},t\right) = \left[-\frac{\hbar^2}{2m}\nabla^2 + V\left(\mathbf{r}\right)\right]\psi\left(\mathbf{r},t\right) = i\hbar\frac{\partial\psi}{\partial t}\left(\mathbf{r},t\right)$$

Interference: ψ = complex function

Interpretation (Bohr, 1927):

ψ = probability amplitude

$|\psi|^2$ = probability



Electron wave functions in hydrogen atom
('standing 3-dim waves')

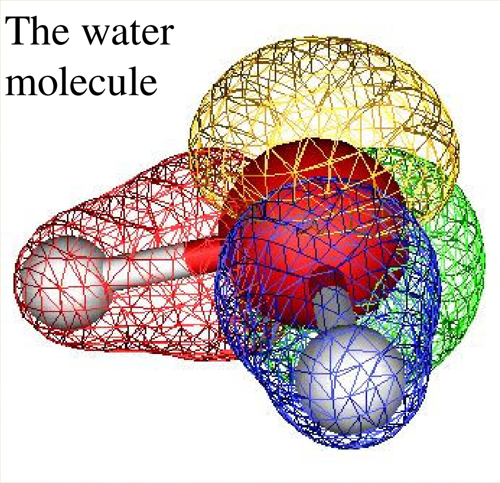
PARTICLE SPECTRUM

1928

Quantum physics explained the existence of 'structure' in nature

Linus Pauling (1928)

The nature of chemical bonds



Atoms, Molecules and the origin of structure were understood.

And the atomic nucleus? Not much progress between 1911 - 1932.

PARTICLE SPECTRUM

n

1932

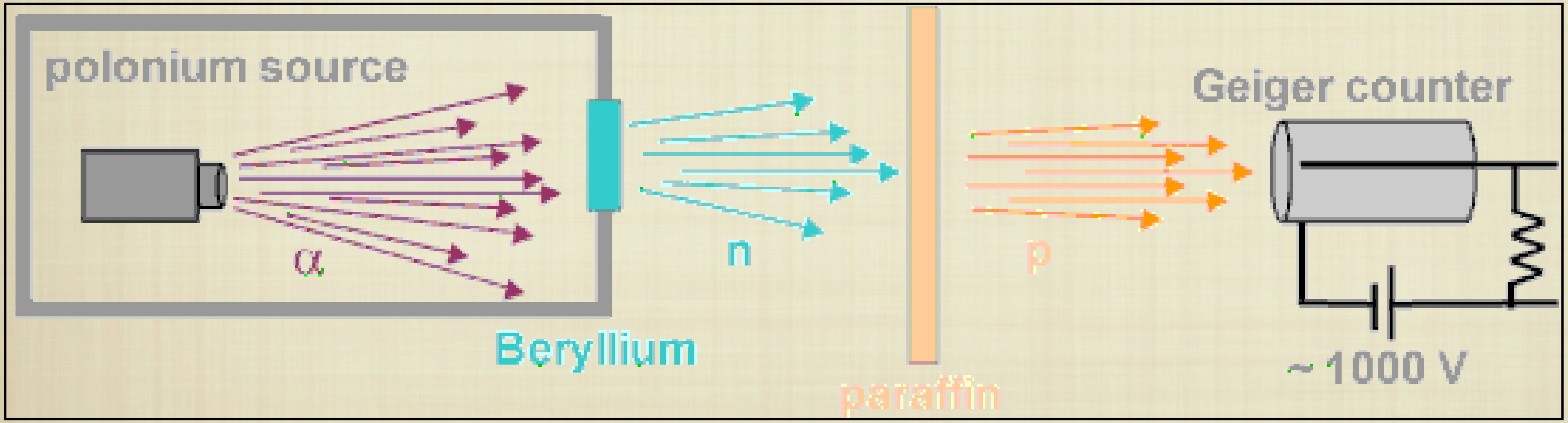
What is the nucleus made of ?

For example: He-4 has only Z=2; what are the other two units of mass due to ?

Heisenberg: Protons and electrons (4 protons and 2 electrons)?

Did not work - the uncertainty relation forbids the presence of electrons in the nucleus!

Chadwick (1932): The neutron

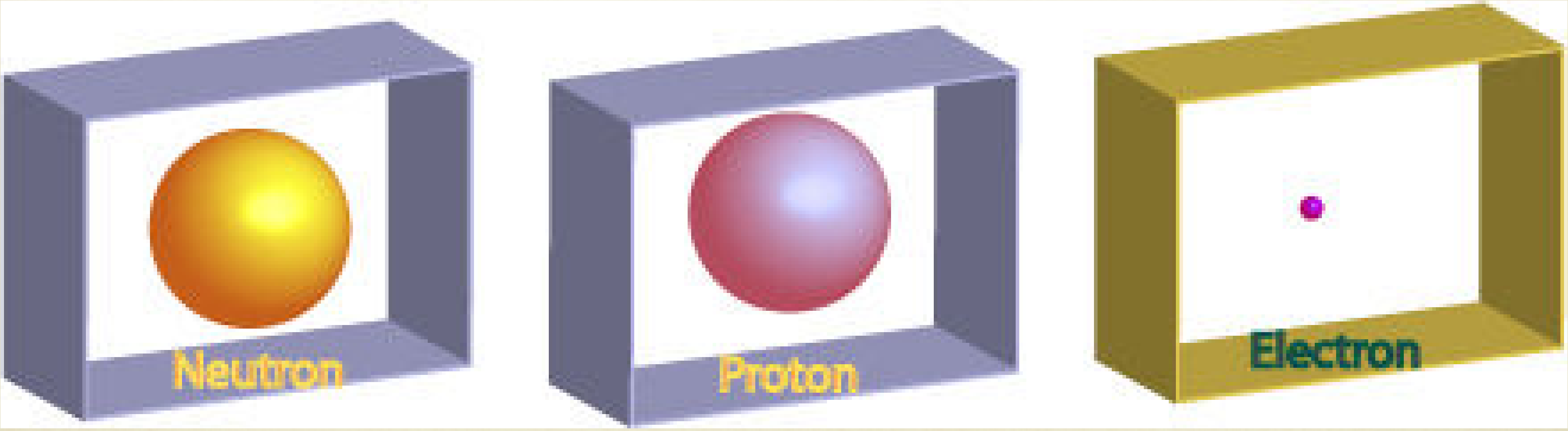


From kinematics: Mass of neutron ~ mass of proton

What keeps everything together? Strong short-range interaction?

PARTICLE SPECTRUM

Fundamental particle spectrum (1932)



Simple, easy to remember
Still taught at schools
(usually in chemistry lessons)

2

Fields

1900

What holds atoms and nuclei together?

1900: two fundamental interactions were known:

$$F_G = G\, m_1\, m_2 \cdot \frac{1}{r^2}$$

$$F_C = Q_1\, Q_2 \cdot \frac{1}{r^2}$$

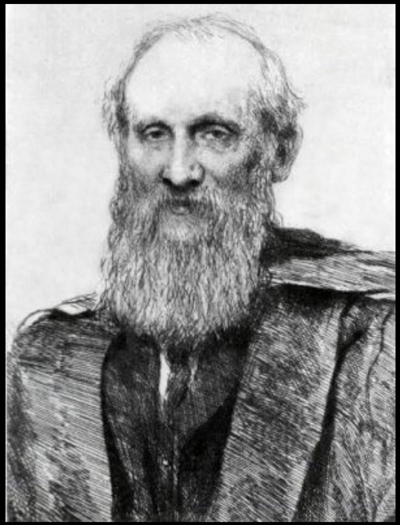
Gravitation

Electromagnetism

Similarities: both have inverse square dependence on radius

Differences: the strength of the forces is vastly different (38 orders of magnitude!)

Remember: in 1900, there were two ‘clouds’ on the horizon of physics:



William Thomson
(Lord Kelvin)

Two clouds:

1) Blackbody radiation
2) Michelson-Morley experiment

Their understanding would lead to

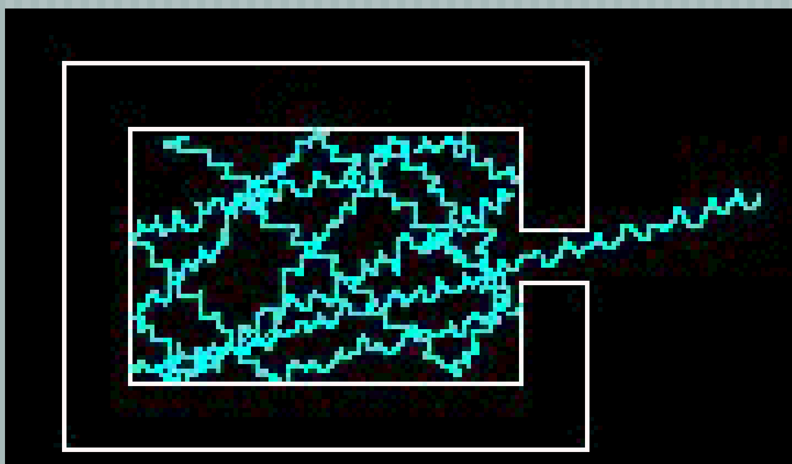
- **quantum theory**
- **relativity**

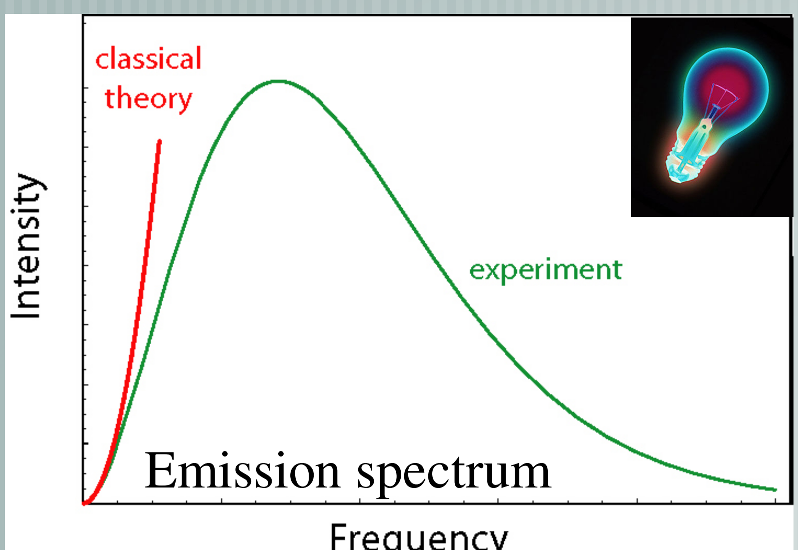
Fields

Photon

'Electromagnetic' interaction

Blackbody radiation





"Black body" absorbs all incoming light;
re-emits thermal equilibrium radiation

"Radiation function" = f(T) only

$$I(\nu) \sim \nu^2 \langle E \rangle$$

average energy of oscillators
(proportional to temperature)

Ok for 'low' temperatures (Jeans law)

Fields

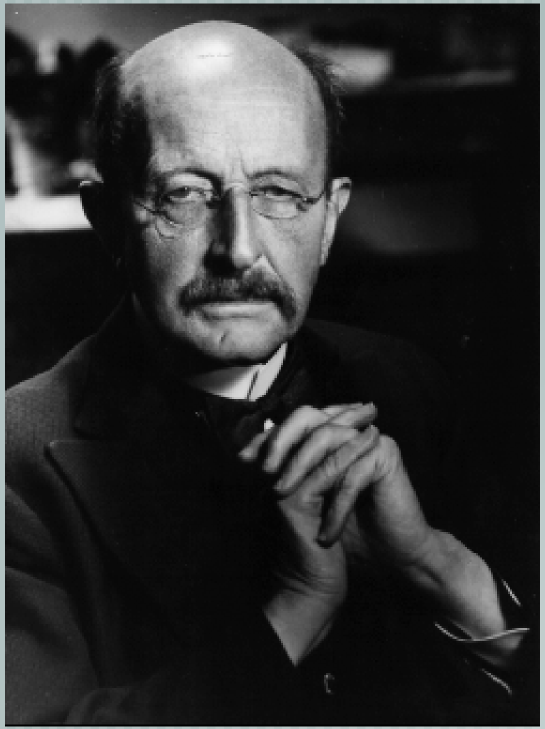
Photon

'Electromagnetic' interaction

An "Act of Desperation"

Oscillators (in the wall of the black body) **emit** ' finite energy elements ' **$\varepsilon = h\nu$**

14 December 1900



Max Planck

Higher frequency means bigger chunks, so it is less likely to find $E \gg kT$

average energy of oscillators

$$I(\nu) \sim \nu^2 \frac{h\nu}{e^{\frac{h\nu}{kT}} - 1}$$

h = new fundamental constant

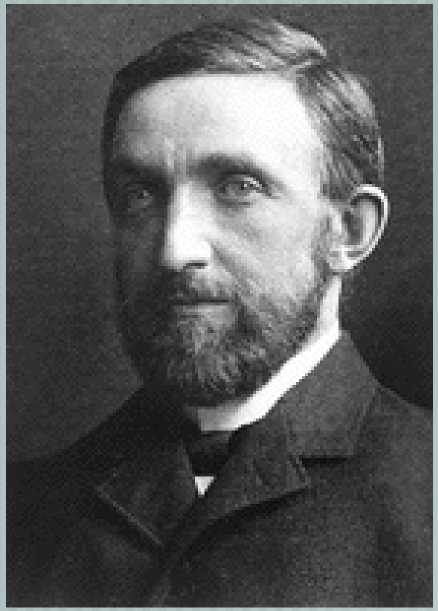
Fields

Photon

'Electromagnetic' interaction

1902

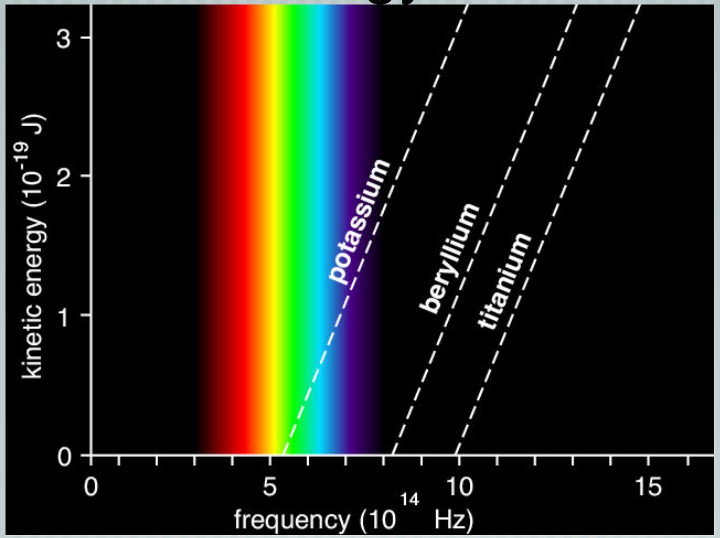
The photoelectric effect



Philipp von Lenard

Cathode rays (electrons) are produced by shining light on metal surfaces.

Classical expectation: Energy of light proportional to square of its amplitude ~ electron energy



Energy proportional to light **frequency** (slope = "h")

"The electron energy does not show the slightest dependence on the light intensity"

Fields

Photon

'Electromagnetic' interaction

"My only revolutionary contribution"

17 March 1905

Albert Einstein

Light is emitted and **absorbed** in **quanta**

$$E_{max} = h\nu - W$$

"A light quantum gives all its energy to a single electron."

Photons are particles.

(Compton,1917, proved it)

Special relativity

Einstein had thought about the ‘medium’ for electromagnetic waves and concluded that there was none.

But how could the speed of light be the same in all inertial frames?

His postulates:

1) Speed of light = constant = c (in vacuum)
2) all inertial frames are equivalent (“relativity principle”)

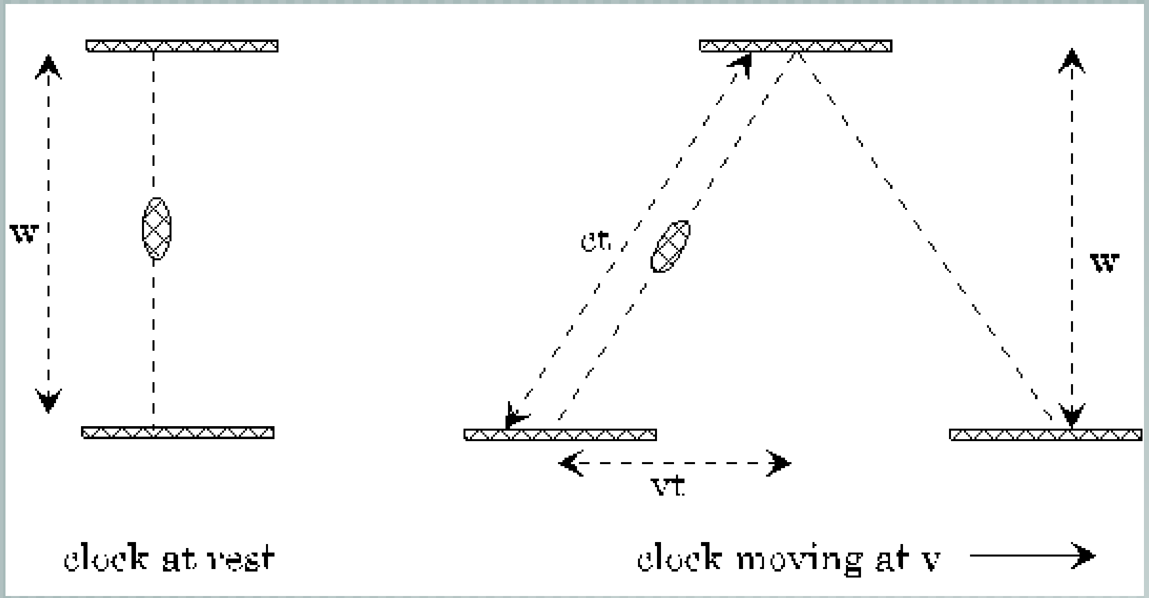
His conclusions:

Since c = constant and speed = (space interval/time interval) -->

space and time cannot be absolute!

Fields

Special relativity



$c^2t^2 = v^2t^2 + w^2$

$t^2(c^2 - v^2) = w^2$

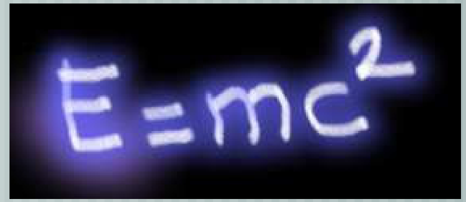
$$t = \frac{w/c}{\sqrt{1-\frac{v^2}{c^2}}} = \gamma \cdot \tau$$

t = time observed for moving frame

τ = time in moving frame (w=c τ)

1) Time dilation, space contraction

2) Modification of Newton's laws, relativistic mass increase.



Fields

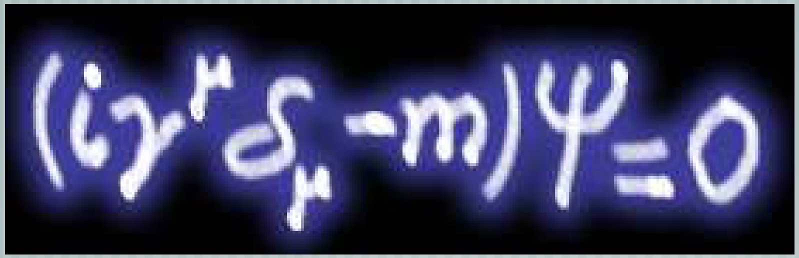
'Electromagnetic' interaction

How could special relativity and quantum physics be united ?

Paul A.M. Dirac
(1928)

$$E^2 = p^2 + m^2 \rightarrow$$

$$E = \pm(\alpha \cdot p) + \beta m$$



$$\Psi = \begin{pmatrix} e^- \uparrow \\ e^- \downarrow \\ e^+ \uparrow \\ e^+ \downarrow \end{pmatrix}$$

Spin

Antimatter

Compare with (non-relativistic)
Schrödinger equation

$$E = \frac{p^2}{2m} \rightarrow \quad i\hbar\frac{\partial}{\partial t}\psi = -\frac{\hbar^2}{2m}\nabla^2\psi$$

CONSEQUENCES:

ELECTRON SPIN EXPLAINED
ANTIPARTICLES MUST EXIST !
ELECTRONS OBEY 'PAULI PRINCIPLE' (1940) - FERMIONS

Fields

'Electromagnetic' interaction

Two crucial (theoretical) predictions by Dirac

The wave function has 4 components (two spin 1/2 particles)

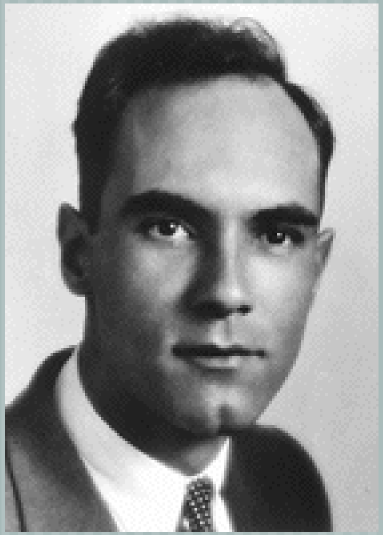
2 components for particle - and 2 components for antiparticle!

Every particle has an antiparticle !

Fields

e^+

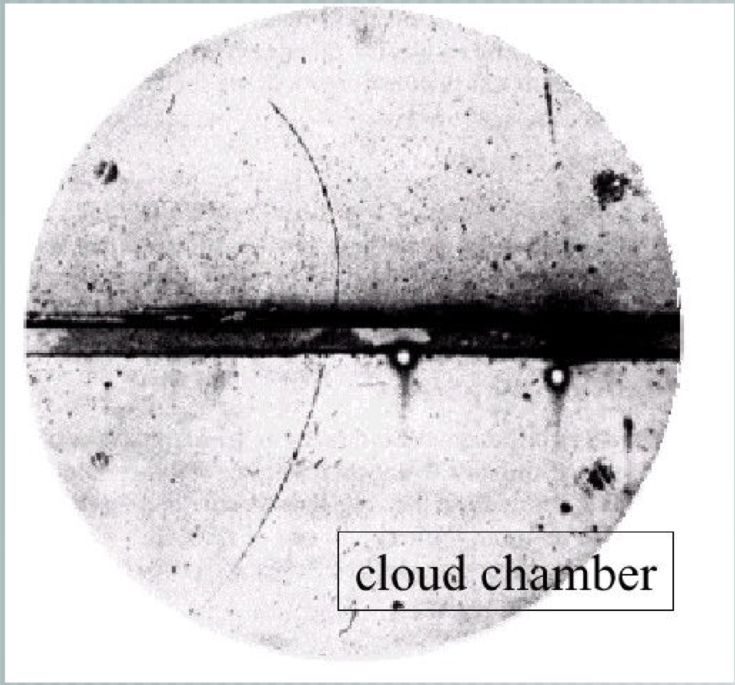
'Electromagnetic' interaction

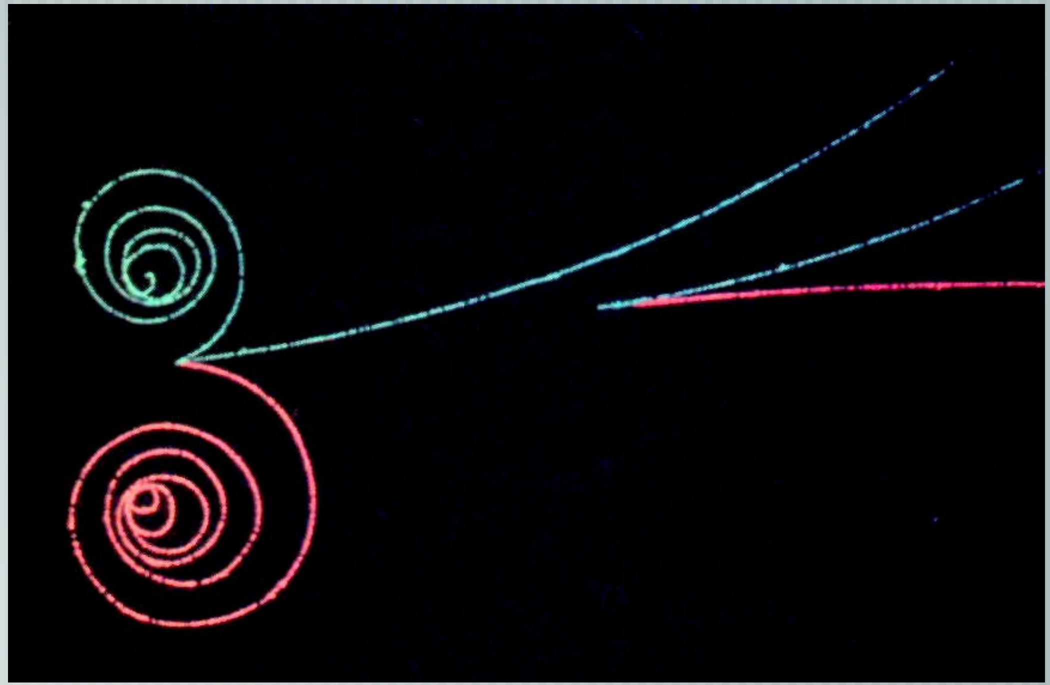


Discovery of the positron

Dirac was right!

Anderson (1932)





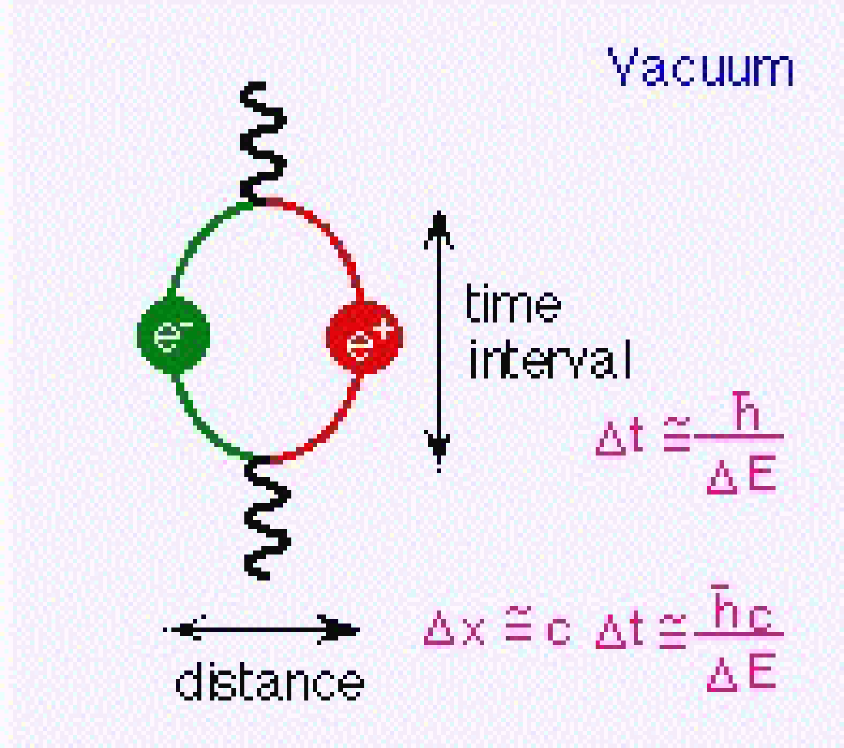
Fields

'Electromagnetic' interaction

NOW THE VACUUM HAD BECOME REALLY MESSY

Quantum physics says that 'oscillators' (e.g. field quanta) cannot be at absolute rest (uncertainty relation)

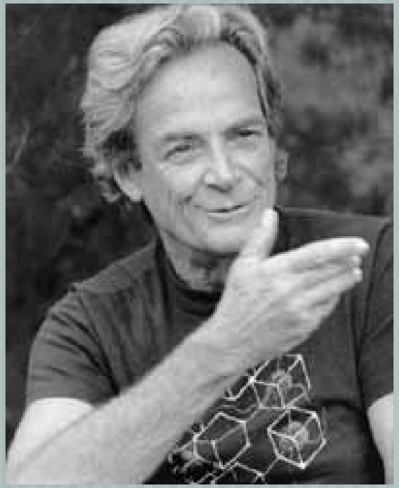
The lowest energy states of e.g. electromagnetic fields can produce (virtual) electron-positron pairs: VACUUM FLUCTUATIONS



Fields

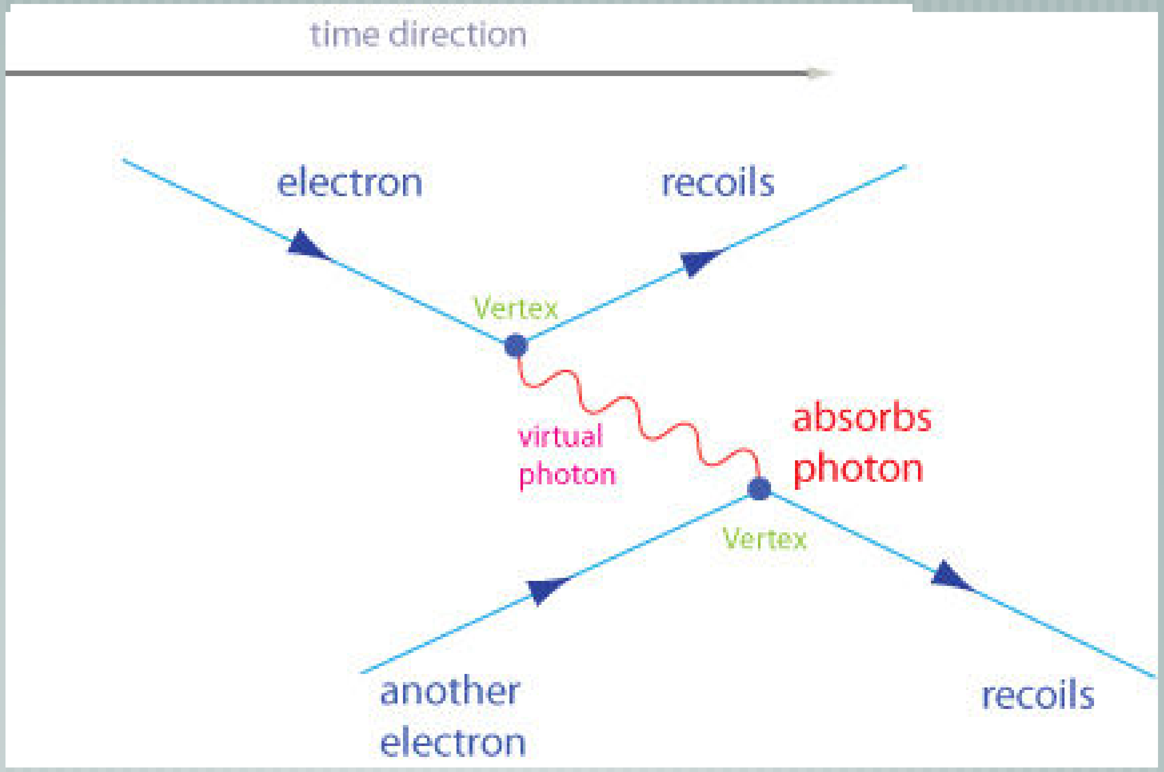
1934 - 1948

How to calculate the interaction of photons and electrons?



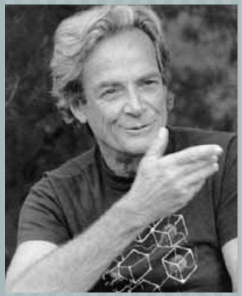
Quantum Electrodynamics

Feynman, Tomonaga, Schwinger



Fields

1934 - 1948

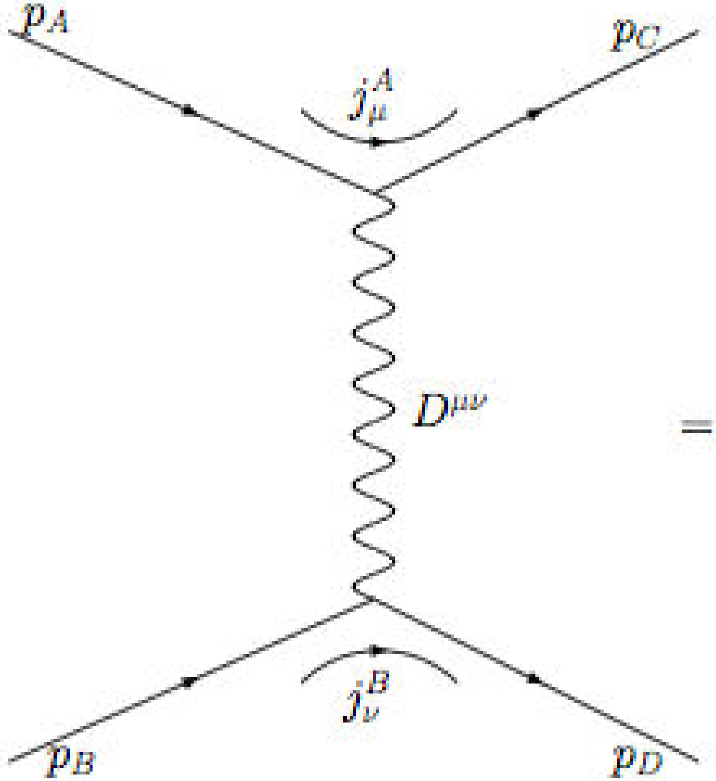


R. P. Feynman

Feynman diagrams

Precise computation rules - in graphical form

Feynman diagrams later became a graphical way to represent all kinds of particle interactions



$$= -i \int \mathrm{d}^4x \, \mathrm{d}^4y \, j_\mu^A(x) \, D^{\mu\nu}(x,y) \, j_\nu^B(y)$$

Fields

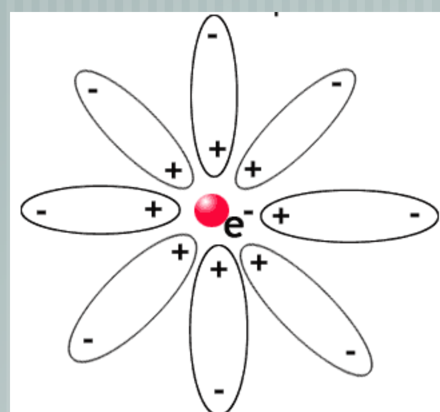
‘Electromagnetic’ interaction

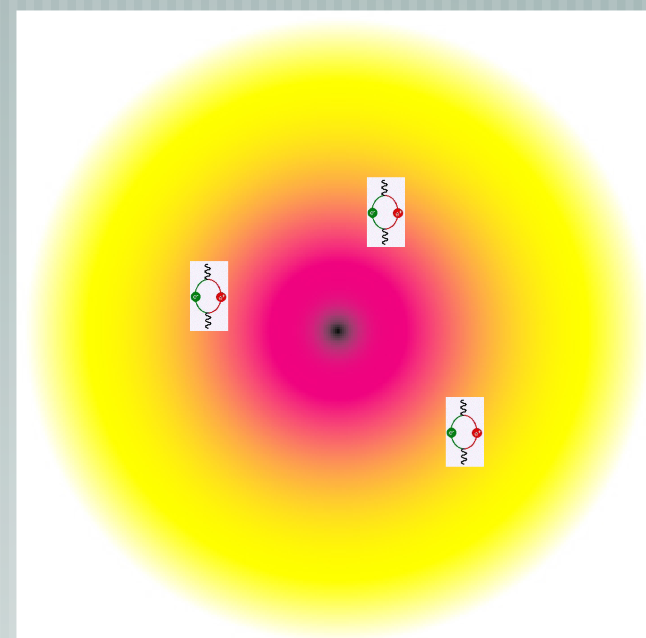
“Renormalization”

The ‘naked’ electron + vacuum fluctuations = measured electron

(“infinite” - “infinite” = “finite”)

the **“dressed” electron**:





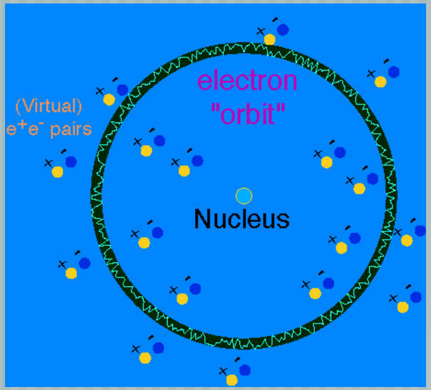
vacuum fluctuations modify its charge and mass (‘Debye shielding’)

Fields

1948

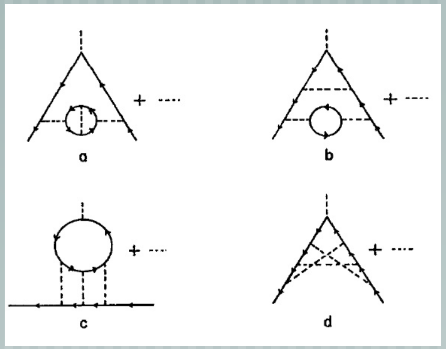
Vacuum fluctuations have observable effects

… and Quantum Electrodynamics allowed to calculate them precisely



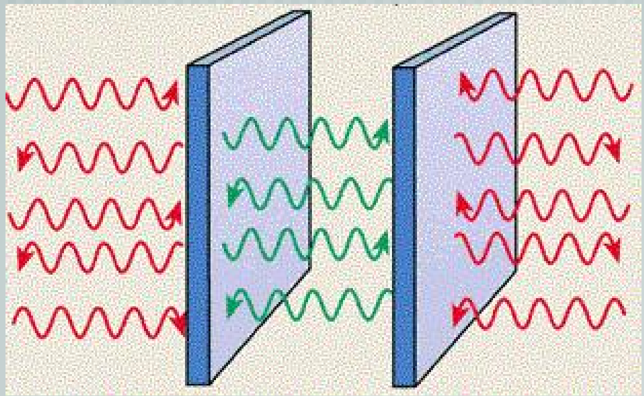
Lamb Shift

(shift of atomic energy levels)



Electron (anomalous) magnetic moment

$$\frac{1}{2}(g-2) = \frac{1}{2}\frac{\alpha}{\pi} - 0.32848\left(\frac{\alpha}{\pi}\right)^2 + (1.183 \pm 0.011)\left(\frac{\alpha}{\pi}\right)^3.$$



Casimir effect

(force on two uncharged metal plates)

Fields

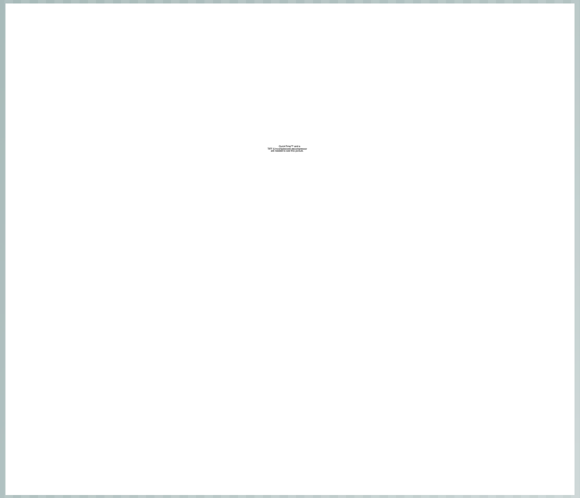
'Electromagnetic' interaction

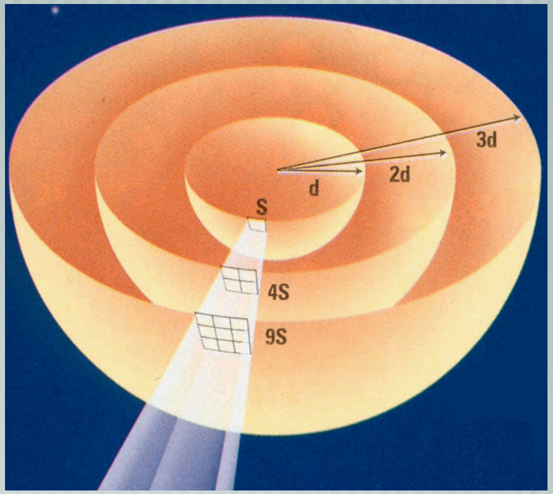
QED: Charged particles interact by exchanging photons

1) **Massless virtual photons are continuously emitted by electric charges**

2) The **$1/r^2$ law** comes from the probability to hit another particle at distance r

(directly connected with the 3 dimensions of space)





$1/r^2$ law

Could that become a model for other interactions?